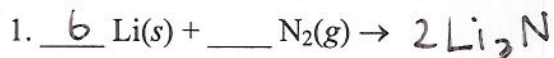
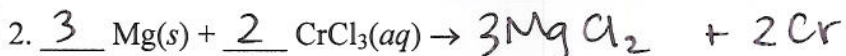


## Types of Reactions – Ch. 8

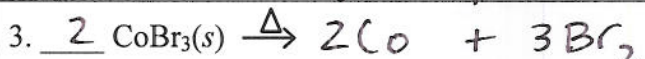
For each of the following reactions, identify the reaction type, predict the products and balance the equation. Include physical states. Word equations must first be converted to formulas.



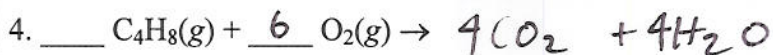
Reaction Type: Synthesis



Reaction Type: single replacement



Reaction Type: Decomposition



Reaction Type: Combustion

5. Aqueous solutions of potassium bromide and silver nitrate react to form a white precipitate.



Reaction Type: double replacement

6. Solid nickel is added to an aqueous solution of iron(II) sulfate. (assume  $\text{Ni}^{+2}$ )



Reaction Type: Single replacement