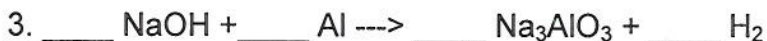


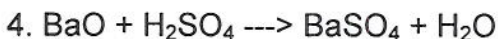
How many grams of O₂ are required to produce 44.0 g of CO₂? (balance first) **52g O₂**

2. Magnesium reacts with hydrogen chloride to produce hydrogen gas and magnesium chloride

How many molecules of magnesium must react to yield 100.0 g of hydrogen gas? **3.01 × 10²⁵ molec Mg**



How many grams aluminum are required to produce 17.5 L of hydrogen? **14.06 g Al**



How much BaSO₄ can be formed from 196.0 g of H₂SO₄?

5. Iron (III) chloride reacts with potassium hydroxide. If 25 g of each reactant are used, what is the limiting reactant? **KOH**

6) If 15.0 g of aluminum sulfide and 10.0 g of water are allowed to react according to the following reaction:



a. Determine the limiting reagent (show a calculation) **H₂O**

b. Calculate how many grams of **each** product will be formed from these reagents. **9.44g H₂S, 14.44g Al(OH)₃**

c. Determine the percent yield if 8.92 g H₂S actually form. **94.5%**

7. 50 grams of zinc are placed in a solution which contains 120 grams of copper (I) acetate. How many moles of copper are formed? **.76 moles Cu**

8. 150 grams of calcium hydroxide were reacted with 60 grams of hydrogen sulfate. How many grams of water were produced? **22.04 g H₂O**

9. 200 grams of silver nitrate react with 300 grams of copper (II) sulfate. How many grams of silver sulfate are formed? **184.5 g Ag₂SO₄**